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Answers: 1. How many grams of silver chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate? 2

$$\text{AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2\text{KNO}_3(\text{aq})$$

0.150 L AgNO_3
0.500 moles AgNO_3 1 moles Ag_2CrO_4
331.74 g $\text{Ag}_2\text{CrO}_4 = 12.4$ g Ag_2CrO_4

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1 L 2 moles ...
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Answer Key Key

CHAPTER 9 REVIEW Stoichiometry

SECTION 2 PROBLEMS Write the

answer on the line to the left

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9. Show all your work in the space

provided. 1. 4 Stoichiometry worksheet

answers chapter 9. 5 mol The following

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equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O_2 form if 3.0 mol of KClO_3 are totally ...

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CHEMISTRY 110 . last first ... Answer

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Answer Key 3] 60.5 mL of HNO₃ are required to react with 25.0 mL of a 1.00 M Barium hydroxide solution: $2 \text{ HNO}_3(\text{aq}) + \text{Ba}(\text{OH})_2(\text{aq}) \rightarrow 2 \text{ H}_2\text{O}(\text{l}) + \text{Ba}(\text{NO}_3)_2(\text{aq})$ (BALANCED) Find the Molarity. of the nitric acid solution .

25.0ml Ba(OH)₂ X . 1L 10. 3. ml X. 1L
1.00mol Ba(OH)₂. X 2mol HNO₃ 1mol

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$\text{Ba}(\text{OH})_2 = 0.0500 \text{ mol HNO}_3 \cdot \text{M} \dots$

WORKSHEET 13 Name - Cerritos College
CHM 130 Stoichiometry Worksheet The following flow chart may help you work stoichiometry problems. Remember to pay careful attention to what you are given, and what you are trying to find. 1.

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Fermentation is a complex chemical process of making wine by converting glucose into ethanol and carbon dioxide: $C_6H_{12}O_6(s) \rightarrow 2C_2H_5OH(l) + 2CO_2(g)$

A. Calculate the mass of ethanol produced if ...

CHM 130 Stoichiometry Worksheet

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Stoichiometry Key Practice Worksheet The correct answer is A. $\text{H}_2\text{C}_2\text{O}_2$ is the limiting reagent: $(4.5\text{mol H}_2\text{C}_2\text{O}_4)(2\text{mol MnSO}_4/5\text{mol H}_2\text{C}_2\text{O}_4)$ According to the equation $2\text{Na(s)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{H}_2\text{(g)}$, sodium reacts with water. Stoichiometry Flashcards | Quizlet Worksheet 14 1 Worksheet #14 Limiting

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